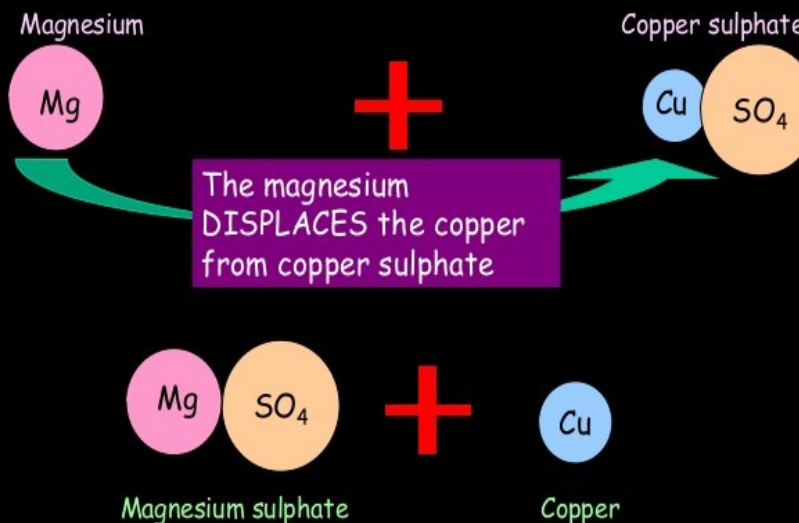


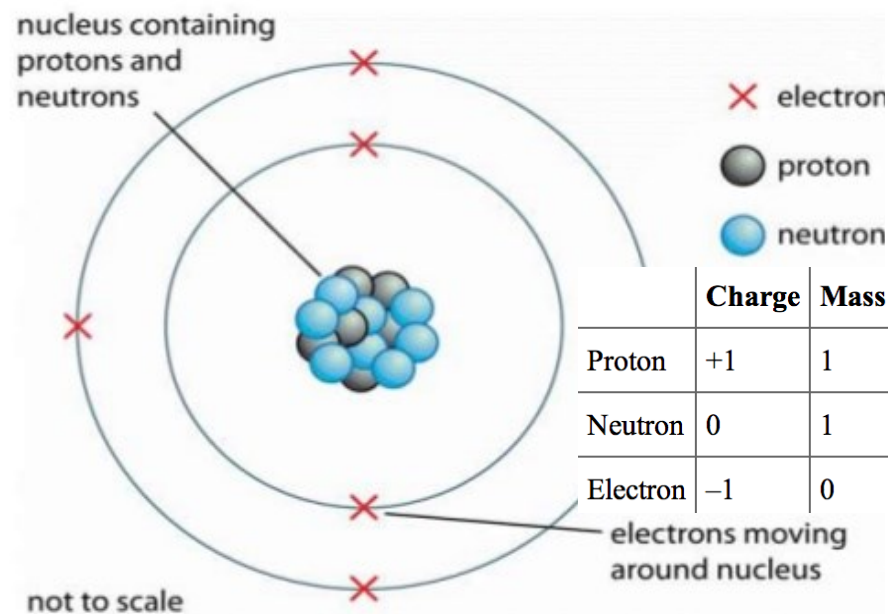
Metals and non-metals Year 8

End point 5 : Predict the outcome of some simple displacement reactions

A displacement reaction is one where a **MORE REACTIVE** metal will **DISPLACE** a **LESS REACTIVE** metal from a compound.



End point 2 : Recall the structure of the atom including the sub-atomic particle mass and charge



End points 3,4 and 6 : Recall the reactivity series of metals

WEIGHT (Light) ↓ (Heavy)	Metal		Reaction with AIR	Reaction with WATER	Reaction with ACIDS
	Potassium	K	Burn vigorously to form metal oxides	React with cold water H₂O (l) to form H ₂ (g) and (metal)OH (aq)	Strong reaction with diluted acid (aq) to form H ₂ (g). Metal replaces H in compound to form a salt.
Sodium	Na				
Calcium	Ca	Burn with decreasing vigour down the series to form metal oxides	Only reacts with steam H₂O(g) to form H ₂ (g) and metal oxide		
Magnesium	Mg				
Aluminium	Al				
Zinc	Zn				
Iron	Fe				
Lead	Pb	React slowly (when heated) to form an oxide layer	No reaction	React with concentrated acid (l) . Metal replaces H to make a salt. Some of the acid decomposes into NO₂ (g) and H₂O (l) .	
Copper	Cu				
Mercury	Hg				
Silver	Ag	No reaction	No reaction	No reaction	
Gold	Au				