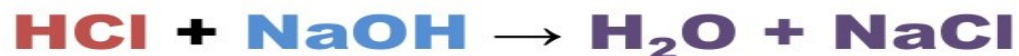
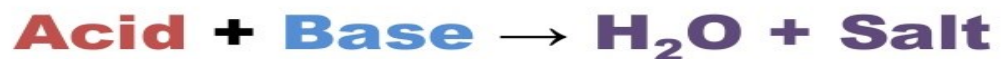


Acids and Alkalis Year 7

End point 3 and 4 : Understand that mixing an acid and an alkali will produce a neutral solution

Acids and Bases Neutralize Each Other

- General formula for acid base reaction



- Salt means any ionic compound formed from an acid/base reaction

NOT JUST NaCl !!

Sodium chloride is produced by the reaction between hydrochloric acid and sodium hydroxide.



Which acid and which alkali would be needed to make the salt potassium chloride?



Acids and Alkalis

Acids are a group of chemicals that contain a H^+ ion example's of which are vinegar, Hydrochloric acid and Sulphuric acid.

Alkalis are a group of chemicals that contain the OH^- ion and have a soapy feel. An example is Sodium Hydroxide. In solid form they are called bases and in solution alkalis.

An acid and alkali can neutralise each other to make a neutral substance.

End point 3 and 4: Be able to name simple salts

Review: Naming Salts

The first part of the salt come from the **metal**.
The second part of the salt comes from the **acid**.

Acid	Salt Formed
hydrochloric acid	- chloride
sulphuric acid	- sulphate
nitric acid	- nitrate

Magnesium + Hydrochloric Acid \rightarrow Magnesium Chloride + Hydrogen

Indicators

End point 1 and 2: Identification of an acid or alkali using the pH scale

Indicators are chemicals that show us whether a substance is an acid or an alkali.

There are many examples including universal indicator, litmus paper and natural ones such as red cabbage.

End point 1: Identification of an acid or alkali using the pH scale

pH scale

